

AMIT AJMANI'S ACADEMY

D-16/82, SECTOR-7, ROHINI

PH: 9810476588, 9953371470

X CHEMISTRY CLASS TEST

M.M.-30

Name of Student: _____ Class & Sub: _____

Topic of Test: Acids, Bases and Salts No. 1 Date : _____

1. Why curd & sour substances are not kept in brass & copper vessels? 1
2. Why do HCl and nitric acid etc. show acidic characters in aqueous solutions, while solutions of compounds like alcohol and glucose do not show acidic character? 1
3. Why does an aqueous solution of an acid conduct electricity? 1
4. Why does dry HCl gas not change the colour of the dry litmus paper? 1
5. While diluting an acid, why is it recommended that the acid should be added to water and not water to the acid? 1
6. Name the sodium compound which is used for softening hard water. 1
7. Write the balanced equations of the following:
 - (a) Dil. Sulphuric acid reacts with Zinc granules.
 - (b) Dil Hydrochloric acid reacts with magnesium ribbon.
 - (c) Dil. Sulphuric acid reacts with aluminium powder
 - (d) Dil. Hydrochloric acid reacts with iron fillings. 4
8. Why do acids not show acidic behaviour in the absence of water? 1
9. Equal lengths of magnesium ribbons are taken in test tubes A and B. Hydrochloric acid is added to A and acetic acid is added to B. In which test tube will the fizzing occur more vigorously and why? 2
10. What is a neutralization reaction? Give two examples. 3
11. What are indicators? Name any two acid-base indicators and mention their colour changes. 3
12. What is a universal indicator? State its use. 2
13. Write the equation formed when zinc granules are added to sodium hydroxide solution and the mixture is warmed. 2
14. A baker found that cake prepared by him is hard & small in size. Which ingredient he forgot to add that would have made it fluffy? Explain. 2
15. What is an olfactory indicator? Name two olfactory indicators. 2
16. Write the natural source from which following acids may be obtained:
 - (a) Acetic acid
 - (b) Citric acid
 - (c) Methanoic acid
 - (d) Lactic acid
 - (e) Tartaric acid
 - (f) Oxalic acid 3

ANSWERS

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M.M.-25

Name of Student: _____ Class & Sub: _____

Topic of Test: Acids, Bases and Salts No. 2 Date : _____

1. Under what soil condition do you think a farmer would treat the soil of his fields with quick lime (Calcium oxide) or slaked lime (Calcium hydroxide) or chalk (Calcium carbonate)? 2
2. What will happen if a solution of sodium hydrogen carbonate is heated? Give the equation of the reaction involved. 2
3. Fresh milk has a pH of 6. How do you think the pH will change as it turns into curd? Explain your answer. 2
4. What is the chemical name and chemical formula of washing soda? 1
5. Explain, how pH change causes tooth decay? How it can be prevented? 2
6. Explain how will you prepare a soda-acid fire extinguisher in laboratory? 3
7. What are amphoteric oxides? Give a few examples. 2
8. Explain pH scale in detail. 3
9. Explain Chlor-alkali process with reaction. 3
10. What happens when bleaching powder reacts with dilute sulphuric acid? Give equation of the reaction involved. 1
11. Why are mineral acids not stored in metal containers? Containers made from which material are safe to store such acids? 2
12. Why does distilled water not conduct electricity whereas rain water does? 2

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M.M.-30

Name of Student: _____ Class & Sub: _____

Topic of Test: Acids, Bases and Salts No. 3 Date : _____

- Q1. What happens when carbon dioxide gas is passed through lime water: a) for a short time, and b) for a considerable time? Write equations. 3
- Q2. What will happen if heating is not controlled while preparing POP? 1
- Q3. Explain four importance of determination of pH value in daily life. 4
- Q4. A white powdery substance having strong smell of chlorine is used for making for disinfecting drinking water supply at water works. Identify the substance. Give its chemical name & write reaction for its preparation. 3
- Q5. Explain what happens when copper sulphate crystals are heated strongly? 2
- Q6. A farmer has found that the pH of soil in his fields is 4.2. Name any two chemical materials which he can mix with the soil to adjust its pH. 1
- Q7. Describe how washing soda is produced starting from sodium chloride (common salt). Write equations of all the reactions involved. 3
- Q8. Explain the action of baking powder in the making of cake (or bread). Write equation of the reaction involved. 2
- Q9. A white powdered solid 'X' when added to water, produces hissing sound. Identify 'X'. How does this compound react with moist hydrogen chloride gas? Write chemical equation. Give some uses of the same compound. 3
- Q10. What are antacids? Give two examples. 2
- Q11. Give the chemical name and chemical formula of Gypsum. 1
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ANSWERS