

AMIT AJMANI'S ACADEMY

D-16/82, SECTOR-7, ROHINI

PH: 9810476588, 9953371470

X CHEMISTRY CLASS TEST

M.M.-30

Name of Student: _____ Class & Sub: _____

Topic of Test: **Chemical Reactions and Equations No. 1** Date : _____

1. Why should a magnesium ribbon be cleaned before burning in air? 1
2. Write the balanced equation for the following reactions:
 - (a) Barium chloride + Aluminium sulphate → Barium sulphate + Aluminium chloride.
 - (b) Solutions of barium chloride and sodium sulphate in water react to give insoluble barium sulphate and the solution of sodium chloride.
 - (c) Hydrogen sulphide gas burns in air to give water and sulphur dioxide. 3
3. Explain with reaction why colour of copper sulphate solution changes when an iron nail is dipped in it? 2
4. What are endothermic and exothermic reactions? Give an example of each. 3
5. A shiny brown coloured element 'X' on heating in air becomes black in colour. Name the element 'X' and the black coloured compound formed with reaction. 2
6. Oil and fat containing food items are flushed with nitrogen. Why? 1
7. Write the products and equation involved when ferrous sulphate is heated. 1
8. How will you know that a chemical reaction has occurred? 2
9. Define the term electrolytic decomposition. Give two examples with reactions. 2
10. Explain with example, the precipitation reaction. 2
11. What is meant by rusting of iron? How is it caused? Explain with equation. 3
12. Give **balanced** chemical equations to show what happens when:
 - (a) Manganese dioxide is reacted with HCl
 - (b) Zinc granules are treated with sulphuric acid.
 - (c) Steam is passed over iron.
 - (d) Carbon dioxide is reacted with water in presence of sunlight.
 - (e) Quick lime is treated with water.
 - (f) Silver chloride is exposed to sunlight.
 - (g) Lead nitrate is heated strongly.
 - (h) Zinc is added to silver nitrate solution. 8

ANSWERS

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X CHEMISTRY CLASS TEST

M.M.-25

Name of Student: _____ Class & Sub: _____

Topic of Test: Chemical Reactions and Equations No. 2 Date : _____

1. Write one equation each for decomposition reactions where energy supplied in the form of heat, light, or electricity. 3
 2. Explain with chemical equation why silver turns grayish in sunlight? 2
 3. What are photochemical reactions? Give two eg. with their chemical equations. 3
 4. A solution of a substance 'X' is used for white washing. Name the substance 'X' and write its formula. Write the reaction of the substance 'X' with water. 2
 5. What happens when carbon dioxide is passed through a freshly prepared slaked lime? 2
 6. Write the balanced chemical reaction when ferrous sulphate is heated. 2
 7. Explain why is respiration considered to be exothermic reaction. 1
 8. Write the products and equation involved when Lead Nitrate is heated. 1
 9. Write the balanced chemical equation involved for the reaction of mixing a solution of lead (II) nitrate and potassium iodide. What is the colour of the precipitate formed? Name the compound precipitated. Is this also an example of double displacement reaction? 3
 10. State a few methods of prevention of rusting? 2
 11. What is slaked lime? How is it formed? 2
 12. What types of reactions are represented by the following equation:
 - (a) $X + Y \rightarrow X^+ + Y^-$
 - (b) $X + YZ \rightarrow XY + Z$
 - (c) $X \rightarrow Y + Z$
 - (d) $X + Z \rightarrow ZX$ 2
 13. Identify the substances oxidized and the substances reduced in the following:
 - (a) $4\text{Na (s)} + \text{O}_2\text{(g)} \rightarrow 2\text{Na}_2\text{O (s)}$
 - (b) $\text{CuO (s)} + \text{H}_2\text{(g)} \rightarrow \text{Cu(s)} + \text{H}_2\text{O (l)}$ 2
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ANSWERS